**Dilute, Concentrated and Saturated Solutions**

Read this information which describes what happens when various amounts of potassium dichromate, an orange chemical, are added to water.



*A solution is* ***dilute*** *if there is a small amount of solid dissolved in a given amount of water.*

If we add one or two crystals of potassium dichromate to a small amount of water and stir well until it dissolves, we will make a weak, or dilute solution.

*A* ***concentrated*** *solution has a large amount of solid dissolved in a given amount of water.*



If we add several more crystals and stir the solution until these dissolve, we get a stronger or more concentrated solution.

**Example:** dilute cordial will have a smaller number of coloured particles so it will be lighter in colour than concentrated cordial.

Someone who has a cup of tea with half a teaspoon of sugar is drinking a more dilute form of sugary tea than a person who adds five teaspoons.



*A* ***saturated*** *solution is a solution where no more solid can dissolve in the given amount of water.*

If we keep adding potassium dichromate a little at a time we eventually reach a time when no more will dissolve. We will now have solid potassium dichromate in the bottom of the beaker and a saturated solution in the beaker.